How do you know if it is a redox reaction

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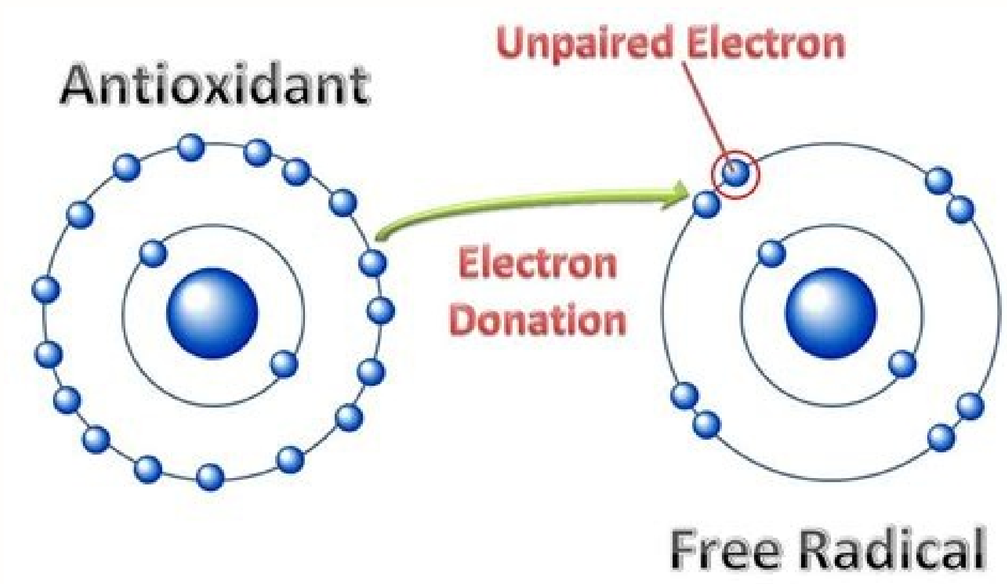
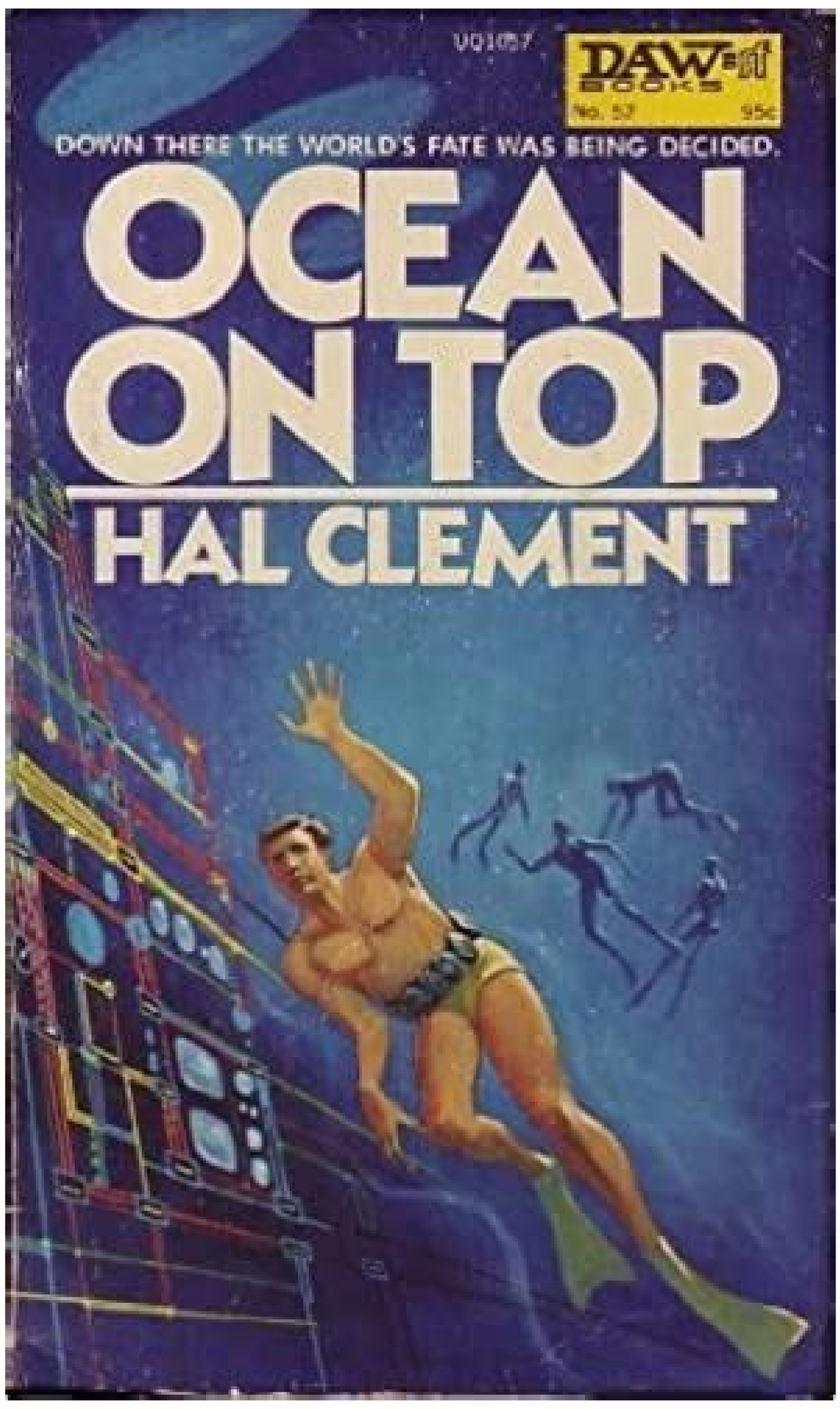
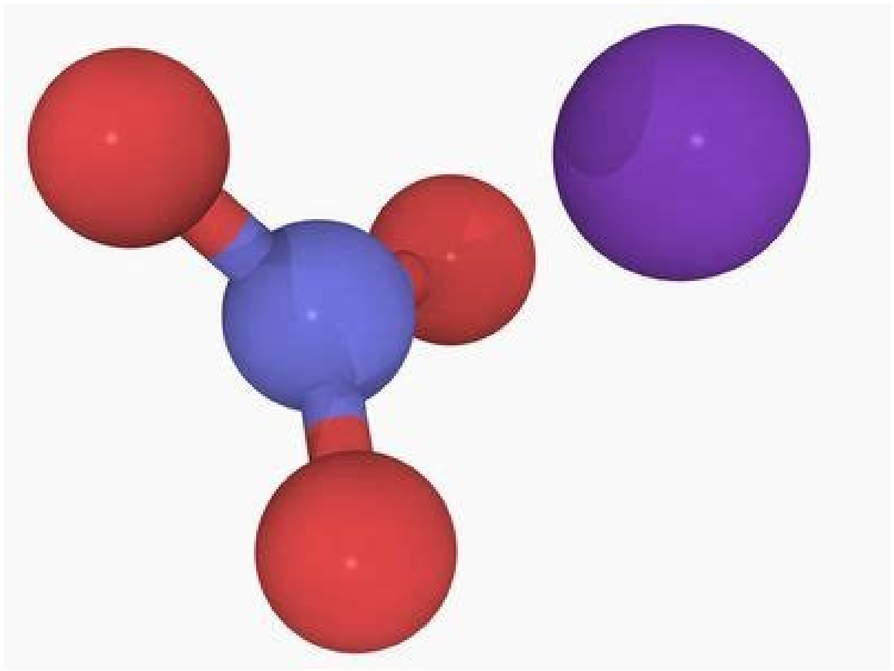
Example 2

- **Next example:**
  $Cu_{(s)} + AgNO_{3\,(aq)} \rightarrow Cu(NO_3)_{2\,(aq)} + Ag_{(s)}$
- Identify the elements undergoing oxidation (Cu) and reduction (Ag). The nitrate group (NO_3) is a spectator ion which we will not include in our half-reactions.
- **oxidation**
  $Cu \rightarrow Cu^{+2} + 2\ e^-$
- **reduction**
  $Ag^+ + 1\ e^- \rightarrow Ag$







What makes a redox reaction. How to know if a redox reaction will occur. Signs of a redox reaction. What does a redox reaction look like. Is it a redox reaction.

If you are seeing this message, it means that we are having problems loading external resources on our website. If you are behind a web filter, make sure that domains * .kastatic.org and * .kasandbox.org are unlocked. Question: A magnesium metal strip is placed in an aqueous 1 mol copper sulfate solution L-1 (ii). Will a spontaneous redox reaction be produced? What do you ask for the question? Determine if the given redox reaction is spontaneous. What data (information) have given you in the question? Magnesium in its standard state, that is, as metal: mg (s) Standard aqueous solution of copper sulfate (II): CUSO4 (AQ) contains copper ions (II), CU2 + (AQ), and ions of sulfate, so42- (AQ) What is there between what you know and what you need to find out? (A) Write The Half-Equations for The Reaction as Given in the Question: Oxidation of Magnesium Metal: mg (s) Ã ¢ Â † Â 'MG2 + (AQ) + 2E-REDUCTION OF COPPER (II) IONS: CU2 + (AQ + 2E- Ã ¢ â † Â 'Cu (s) (B) Use to table of Standard Reduction Potentials to find the value of EO for Each Half-Reaction: mg (s) Ã ¢ Â † Â' mg2 + (AQ) + 2E- Ã, EO = +2.36 V CU2 + (AQ) + 2E- Ã ¢ Â † Â 'Cu (s) Ã, EO = +0.34 V (c) Add The Equitions Together, and Add The EO Value for Each Half -Reaction Together Mg (s) Ã ¢ Â † Â 'Mg2 + (AQ) + 2E- Ã, EO = +2.36 V CU2 + (AQ) + 2E- Ã ¢ Â † Â' Cu (s) Ã, EO = +0.34 V mg (s) + CU2 + (AQ) Ã ¢ Â † Â 'Mg2 + (AQ) + Cu (s) Ã, EO = +2.36 + +0.34 EO = +2.70 V Decide if The Redox Reaction as Written is spontaneous: (A) Spontaneous Reaction IF EO (Redox)> 0 (B) Not at Spontaneous Reaction IF EO (Redox) 0 The Redox Reaction Is Spontaneous Is Your Answer Plausible? Consider the activity (reactivity) of the metals involved: (a) magnesium (mg), a metal of group 2, is a very active metal (reactive) (b) copper (cu), is a transition metal and is Much less active (reagent) than magnesium metal. Therefore, we hope that magnesium exists in the form of ions in soluces. This means that we hope that the copper exists as The spontaneous reaction would be for to magnesium ions and for copper ions to accept these electrons to produce copper metal: that is, the spontaneous reaction would be: Cu2+(aq) + Mg(s) → Mg2+(aq) + Cu(s) And this reaction is the same as the reaction they gave us, that is, they gave us a piece of magnesium metal, Mg(s), to place in an aqueous solution of copper ions (II), Cu2+(aq), so we hope that a reaction between these two substances will occur spontaneously! Declare your solution to the problem "is the redox reaction given spontaneously or not": The Mg(s) + Cu2+(aq) → Cu(s) + Mg2+(aq) reaction will proceed spontaneously in the front direction. Can you apply this? Join AUS-e-TUTE! Take the test now! If we know the mass, pressure, volume and temperature of a gas, we can calculate its molar mass using the ideal gas equation. Remember that the ideal gas equation is given as: PV = NRT. We can reorganize this equation in terms of molars (n) and then solve for its value. Once we have the value for the molars, we can divide the gas mass by the molars to get the molar dough. Here is how to do it: Using the ideal gas equation to calculate the molar mass We now apply the equation (3) and (4) to solve the following problem: Mary filled her metallic cylinder of 5.90 L with 1066 g of gas at a pressure of 2025 psi and at a temperature of 25 °C. calculates the molar mass of this gas and uses the value of the molar mass to identify the gas. Use the following additional information to help you calculate the correct answer: 1 atm = 14.7 psi and R = 0.08206 L. atm/(mol.K). Strategy To calculate the molar mass of the gas, we must first calculate the moles of the gas. To do this, we must use the equation (3); n = PV/RT After we get our moles, then we will replace this value and mass in equation (4) to resolve for the molar mass Units must be consistent. From units in the constant of universal gas, can bethat the volume is in liters, pressure in am and temperature in Kelvin. This means that we must convert the units of pressure to pressure ATM ATM psi and Celsius grade temperature units to kelvin. to convert from the Celsius grade to kelvin, we simply add 273 to our temperature in celsius. i.e. 273 + 25 ° c = 298 k. to convert 2025 psi to atm, we simply multiply 2050 psi by the conversion factor, 1 atm / 14.7 psi. i.e.: the pressure of the psi to the atm solution that calculates the molar mass of the value of the molar mass, may say that this gas is oxygen, with the chemical formula o2. How did we know? We knew that because of the periodic table, the molar mass of an oxygen atom is approximately 16.00 g/ mol. and since the oxygen molecule consists of two oxygen atoms, then it is deduced that to obtain the molar mass of this molecule, we must multiply 16.00 g/ mol by 2. and if we do, we will get: 32.00 g/ mol. This value is almost the same as what we got. What are some things to take into account when solving problems of the gas law? here are some tips to take into account when you smell the ideal equation to solve gas problems: make sure the units connected to the variables are consistent with the units in the universal gas constant. for example, our answer in the previous problem would have been wrong if we had heard pressure attached to psi instead of atcule

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